The Coefficient Connection

$NaHCO_3 + HCl \rightarrow NaCl + H_2CO_3$

Calculation #1

$$1.99\ \text{g}\ NaHCO_3 \times \frac{1\ \text{mol}\ NaHCO_3}{84.0\ \text{g}\ NaHCO_3} = 0.024\ \text{mole}\ NaHCO_3$$

The Coefficient Connection

$NaHCO_3 + HCl \rightarrow NaCl + H_2CO_3$

Calculation #2

Mass of dish, glass, and NaCl	=	84.3 g
Mass of dish, glass	=	82.9 g
Mass of NaCl	=	1.36 g

The Coefficient Connection

$NaHCO_3 + HCl \rightarrow NaCl + H_2CO_3$

Calculation #3

$$1.36\ \text{g}\ NaCl \times \frac{1\ \text{mol}\ NaCl}{58.5\ \text{g}\ NaCl} = 0.023\ \text{mole}\ NaCl$$

The Coefficient Connection

$NaHCO_3 + HCl \rightarrow NaCl + H_2CO_3$

Calculation #4

$$\frac{0.023\ \text{moles}\ NaCl}{0.024\ \text{moles}\ NaHCO_3} = \frac{1\ \text{mole}\ NaCl}{1\ \text{mole}\ NaHCO_3}$$

Chem Is Try #2

$NaHCO_3 + HCl \rightarrow NaCl + H_2CO_3$

$$2\ \text{g}\ NaHCO_3 \times \frac{1\ \text{mol}\ NaHCO_3}{84.0\ \text{g}\ NaHCO_3} \times \frac{1\ \text{mol}\ NaCl}{1\ \text{mol}\ NaHCO_3} \times \frac{58.5\ \text{g}\ NaCl}{1\ \text{mol}\ NaCl}$$

1.39 g NaCl